

"Express Mail" mailing label number EV 327 129 208 US.

Date of Deposit October 29, 2003.

Our Case No. 10322/63

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE
APPLICATION FOR UNITED STATES LETTERS PATENT

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TITLE:

WATER PURIFICATION BY
POLYMER COLLOIDS

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WATER PURIFICATION BY POLYMER COLLOIDS

BACKGROUND

[001] The present invention relates to a novel process for the purification of water. Dissolved organic substances typically found in water supplies include man-made compounds and the natural products of plant decay, including humic acid, which can give undesirable color, taste, and odor to water. Dissolved organics also interfere with the water purification process by clogging filters and fouling resin beds. Furthermore, during treatment of the water supply at drinking water treatment facilities, natural organic matter can also react with chemical disinfectants, such as chlorine, to produce chlorinated-organic compounds, many of which are known carcinogens.

[002] Activated carbon has long been used to adsorb organic compounds from water supplies. Typically, carbon filters are used prior to other purification processes to partially purify the water and prevent fouling of ion-exchange beds or membranes. Presently, activated carbon pre-filters are used before high performance water purification membranes, such as reverse osmosis membranes. Without a pre-filter, such membranes would rapidly clog with organic contaminants. However, all activated carbons have the disadvantage of requiring disposal or regeneration after their adsorption capacity is exhausted.

[003] Many activated carbons, including those in powdered form, are never regenerated and remain in the residuals or sludges generated from water treatment. These residuals or sludges must then be disposed of as solid waste, usually in a landfill, which poses a secondary pollution problem.

[004] Unlike their powdered counterparts, granular activated carbons can be regenerated using a high temperature process where the adsorbed

organic materials are volatilized or burned from the carbon. This process is energy intensive, and can cause secondary air and water pollution problems.

BRIEF SUMMARY

[005] In a first aspect, the present invention is a colloid comprising a sulfone polymer.

[006] In a second aspect, the present invention is a method for purifying water, comprising contacting a colloid comprising a polymer with water, the water comprising organic matter; and separating the colloid from the water.

[007] In a third aspect, the present invention is a method of purifying a colloid comprising contacting a colloid further comprising organic matter with an alkali solution.

BRIEF DESCRIPTION OF THE DRAWINGS

[008] Figure 1 shows a representative laboratory scale synthesis apparatus for forming polymer colloids.

[009] Figure 2 shows the adsorption isotherm for humic acid on sulfone polymer colloids and NORIT powdered activated carbon.

[0010] Figure 3 shows the regeneration of sulfone polymer colloids at increased pH and the readsorption of humic acid onto the colloids.

[0011] Figure 4 shows the adsorption kinetics for humic acid at different sulfone polymer colloid concentrations.

DETAILED DESCRIPTION

[0012] The present invention includes using colloids, such as sulfone polymer colloids, to adsorb organic contaminants from water. The colloids

may be made, for example, by dissolving a sulfone polymer in a solvent and precipitating the sulfone polymer as a polymer colloid. When the colloids are added to water containing organic contaminants, the contaminants are adsorbed onto the colloids. Removal of the colloids from the water results in removal of the organic contaminants, purifying the water. The colloids may be regenerated, or cleaned of the adsorbed organic contaminants, by exposing them to base. The regenerated colloids may then be reused.

[0013] The colloids may directly replace powdered activated carbon as it is presently used in drinking water treatment, for taste and odor control and adsorption of natural and synthetic organic pollutants. Rather than being discarded in the water treatment sludge, as is presently done with activated carbon, the colloids may be regenerated and reused. Additionally, the colloids can be used in place of an activated carbon pre-filter. When the colloids are made from the same or similar polymer as the purification membrane, superior pre-filtration is possible.

[0014] As described in Van Nostrand's Encyclopedia of Chemistry, pp. 272-276 (Douglas M. Considine ed., Van Nostrand Reinhold Co. 1984), colloids are disperse systems with at least one particle dimension averaging in the range of 10^{-6} to 10^{-3} mm. Particles may be defined as liquid or solid. Examples include sols (dispersions of solid in liquid), emulsions (dispersions of liquids in liquids), and gels (systems, such as jelly, in which one component provides a sufficient structural framework for rigidity and other components fill the space between the structural units). Preferably, the polymer colloids of the current invention are sols or sol-gels.

[0015] The polymer colloids may be precipitated when a solution containing the polymer is added to a liquid in which the polymer has lower solubility than the solvent of the solution. The solution is formed by dissolving the polymer in a solvent or mixture of solvents that has a higher solubility toward the polymer.

[0016] When the solution containing the polymer is added to a liquid having a lower solubility for the polymer, polymer colloids precipitate. Various solvents, solvent mixtures, surfactants, wetting agents, and acids can be used to tailor the morphology of the colloids. While many large scale production methods could be used, as known to those of skill in the art, a syringe pump is appropriate on the laboratory scale. A representative laboratory scale synthesis apparatus is shown in Figure 1.

[0017] Polymers useful in the present invention include, sulfone homopolymers and copolymers such as polymers of polysulfone, polyethersulfone, polyphenylsulfone, and sulfonated polysulfone; homopolymers and copolymers of cellulose acetate, polyacrylonitrile (PAN), polyetherimide, and poly(vinylidene fluoride) (PVDF); and mixtures thereof. Such polymers may be purchased from AMOCO PERFORMANCE PRODUCTS, INC. (Alpharetta, GA) under the trade names of UDEL (polysulfone), MINDEL (sulfonated polysulfone), RADEL-A (polyethersulfone), and RADEL-R (polyphenylsulfone). They are also available from ALDRICH, Milwaukee, WI.

[0018] Suitable average molecular weights for polysulfone useful in the current invention range preferably from 10,000 to 45,000, more preferably from 17,000 to 35,000, and most preferably from 26,000 to 27,000. Suitable average molecular weight for polyethersulfone useful in the current invention range from 8,000 to 28,000, preferably from 13,000 to 23,000, and most preferably from 16,000 to 20,000. Suitable average molecular weights for poly(vinylidene fluoride) useful in the current invention range preferably from 100,000 to 600,000, more preferably from 180,000 to 534,000, and most preferably from 275,000 to 530,000. Suitable average molecular weights for polyacrylonitrile useful in the current invention range preferably from 30,000 to 150,000, more preferably from 60,000 to 110,000, and most preferably from 80,000 to 90,000. All average molecular weights are weight average molecular weights.

[0019] The solution containing the polymer includes the polymer and one or more solvents in which the polymer demonstrates solubility. Any solvent that permits colloid formation when the polymer solution is added to a liquid in which the polymers have lower solubility may be used. Additionally, the solution may contain an acid. Any acid which is compatible with the selected polymer and solvent system can be used if colloids form when the solution is mixed with a liquid in which the polymer has lower solubility. Although not intending to be bound by any particular theory, it is believed that the acid (a weak Lewis acid) complexes with the solvent (a weak Lewis base) to form a complex which breaks up when mixed with the liquid. Many acid/solvent systems are possible and are more fully described, for example, in Fritzsche, et. al., *Journal of Membrane Science*, **46**, 135 (1989).

[0020] Suitable solvents include *N*-methyl pyrrolidine (NMP), *N,N*-dimethylformamide (DMF), dimethyl sulfoxide (DMSO), acetone, and dioxane, and are available from ALDRICH, Milwaukee, WI. Suitable acids include organic acids, such as propionic acid.

[0021] Additionally, surfactants may be added to the solution to stabilize the colloids and otherwise vary their morphology. While any surfactant, including anionic, cationic, or non-ionic, may be used, preferable surfactants include sodium lauryl sulfate, TRITON X-45, and TRITON X-100, or mixtures thereof. Wetting agents, such as alcohols, may also be added.

[0022] Scanning electron micrographs show that small, relatively uniform spherical particles of colloids with average individual diameters of preferably from 10 to 1000 nanometers (nm), more preferably from 25 to 500 nm, and most preferably from 50 to 100 nm, are formed during precipitation. Aggregates of the colloid particles attain various shapes and sizes with a wide size distribution. Average aggregate diameters are from 10 μm to 1000 μm , preferably from 25 μm to 500 μm , and most preferably from 100 μm to 500 μm .

[0023] When the polymer colloids are added to water containing organic matter, the contaminating organic matter is adsorbed. The addition may be carried out in any appropriate agitated vessel or fluidized reactor. Although variables, including temperature, contaminant concentration, and colloid concentration affect the rate of adsorption, the organic matter is typically adsorbed onto the colloids within minutes to hours.

[0024] Organic matter includes hydrocarbons, hydrophobic pollutants, or pollutants with mixed hydrophobic/hydrophilic properties that pollute water by imparting an undesirable color, taste, or odor, as well as any other carbon-containing compound. Preferably, the present invention removes natural organic matter; natural organic matter includes carbon containing material typically found in drinking water supplies. Although many types of organic matter contaminants may be found in water, humic acid is one of the most common. Other organic matter contaminants include benzene, toluene, proteins, geosmin (a natural organic compound leached from soils), and 2-methylisoborneol (MIB) (a natural organic compound of aquatic biological origin).

[0025] Purified water is formed by removing the colloids that contain the organic contaminants from the water. As shown in Figure 2, the colloidal polymers have a larger adsorption capacity for humic acid than activated carbon adsorbents typically used in the treatment of drinking water, including NORIT powdered activated carbon.

[0026] Any method may be used to remove the contaminated colloids from the water, such as centrifugation, filtration, gravity decantation, counter-current decantation, and packed column filtration. Filtration methods typically entail passing the purified water containing the contaminated colloids through a filter with pore sizes of a smaller average diameter than the average diameter of the colloids.

[0027] Gravity and counter-current decantation and fluidized bed methods rely on the colloids, or the substrates to which they are bound, having a greater density than the purified water. Gravity decantation describes methods where the colloids are allowed to settle due to gravity from the purified water, which is then removed without significantly disturbing the settled colloids. Counter-current decantation and fluidized beds suspend the colloids in the contaminated water stream using fluid movement.

[0028] Similarly, the colloids may be immobilized on a substrate which is then filtered out. Alternatively, the colloids are immobilized on a membrane or a packed column of substrate beads through which the contaminated water passes. Typically, a suspension containing the colloids are vacuum filtered through a membrane, thus trapping the colloids or their aggregates on the membrane. Suitable substrates include any material which is chemically compatible with the colloids and has the ability to bind the colloids while being physically larger, such as glass or plastic beads or mesh. The colloids may also be held in a container, such as a nylon mesh bag, which is permeable to water but not the colloids.

[0029] Once the contaminated colloids are removed from the purified water, they may be regenerated by chemically desorbing the organic contaminants. Unlike activated carbon, this allows reuse of the colloids for continued water purification and optional recovery of the organic contaminants (Figure 3). In many instances, the organic contaminants have value as a fertilizer for plants.

[0030] The organic contaminants are desorbed from the colloids by exposing the colloids to an alkali solution. Although the alkali solution may be of any concentration, it preferably has a free hydroxide concentration of 1×10^{-4} to 10 N, more preferably of 1×10^{-3} to 5 N, and most preferably of 1×10^{-2} to 1 N. Any alkali solution may be used, such as a solution of sodium

hydroxide, ammonium hydroxide, potassium hydroxide, calcium hydroxide, or mixtures thereof. Sodium hydroxide is most preferred.

[0031] Once the colloids are exposed to elevated pH, they are removed from solution using separation methods apparent to those of skill in the art, including those previously described, and are returned to the water source to remove more organic contaminants.

EXAMPLES

Example 1

[0032] Two grams of polysulfone (PS) (26,000-27,000 MW) was added to 56 grams of N-methyl pyrrolidine (NMP) solvent. The mixture was shaken in a gyrator until the polysulfone solids fully dissolved (approximately 1 day). Forty two grams of propionic acid (PA) was added to the PS/NMP solution to form a mixture with a molar ratio of 1:1 PS/NMP to PA. The mixture was shaken until a homogeneous solution formed. The solution was stored at room temperature.

[0033] A syringe was then filled with 2.5-mL of the stored solution and placed in a syringe pump with a pumping rate of 1.5 mL/hr. A beaker was filled with 500 mL of deionized water which was stirred by a magnetic stirrer. The syringe pump was then used to pump the polysulfone containing solution through a capillary tube into the stirred water at about 25° C. Polymer colloids formed in the deionized water. When agitation was stopped, the colloids quickly settled to form aggregates.

[0034] Colloid samples were observed under ordinary optical microscopy (aus JENA Laboval 4) with 200X magnification. Samples were also examined under Environmental Scanning Electron Microscopy (ESEM). The polysulfone colloids were concentrated onto 0.22- μ m nylon filter paper by filtering the colloids solution with a syringe filter. The filter paper was then

dried and coated with a gold-palladium sputter for ESEM analysis. ESEM analysis showed that the colloids are spherical in nature, having a diameter of about 50 nanometers, and are clumped into small aggregates.

Example 2

[0035] To test the performance of the PS colloids for water purification, Suwannee River Humic acid, obtained from the International Humic Substances Society, University of Minnesota, St. Paul, Minnesota was used as a model. Humic acid was adsorbed by polysulfone colloids of varying concentration (100 mg/L, 10 mg/L, and 1 mg/L), but with constant humic acid concentration as follows:

[0036] A 15 mg/L humic acid solution was prepared by dissolving 15 mg of humic acid (Suwannee River Humic acid, International Humic Substances Society) in a 1-Liter Erlenmeyer flask, with 10^{-3} M Na_2HPO_4 buffer. The solution pH was adjusted to 6.5 with 1N HCl.

[0037] A 200 mL aliquot of each colloid composition (100 mg/L, 10 mg/L, and 1 mg/L) was mixed with 100 mL of the 15 mg/L humic acid solution in a glass jar. The jar was covered with a sheet of aluminum foil, and the cap was screwed on tightly over the foil. The glass jars were then put onto a gyrator shaker at approximately 22° C. A 10 mL sample was drawn from each jar on the 1st, 3rd, 5th, and 7th days and passed through a 0.22 μm nylon filter to remove the colloids. The percent of humic acid (HA) remaining in solution was then calculated from UV adsorption data at 254 nanometers. Recorded UV readings were the average of 10 replicates for each colloid concentration and are presented below in Table I.

Time (days)		0	1	3	5	7
Control 5 mg/L HA	UV-ads	0.153	0.151	0.150	0.149	0.150
100 mg/L Colloid 15 mg/L HA	UV-ads	0.153	0.029	0.02	0.0194	0.02
	% HA Adsorbed	0	80.79	86.67	86.98	86.67
100 mg/L Colloid 15 mg/L HA	UV-ads	0.153	0.036	0.025	0.023	0.025
	% HA Adsorbed	0	76.16	83.33	84.56	83.33
10 mg/L Colloid 15 mg/L HA	UV-ads	0.153	0.088	0.081	0.076	0.069
	% HA Adsorbed	0	41.72	46.00	48.99	54.00
1 mg/L Colloid 15 mg/L HA	UV-ads	0.153	0.139	0.131	0.124	0.117
	% HA Adsorbed	0	7.95	12.67	16.78	22.00

Table I

[0038] The adsorption kinetics for humic acid at different polysulfone colloid concentrations are shown in Figure 4. Almost 87% of the humic acid was adsorbed onto the 100 mg/L polysulfone colloids. Two trials of five aliquots each showed good reproducibility. The control samples demonstrated that the adsorption of humic acid on the wall of the glass jar was negligible, and that only polysulfone colloids contribute to the adsorption of humic acid over the 7 day period. Most adsorption occurred during day one and then plateaued. Based on the final sample taken on the 7th day, an isotherm was determined as shown in Figure 2. Using the Freundlich isotherm to fit the data, a $F = 0.103$ (mg HA/mg polysulfone) and a $1/n = 1.9224$ were obtained.

Example 3

[0039] In order to check whether the polysulfone colloids are reusable, a set of adsorption, desorption and readsorption tests were performed with 300 mL aliquots of 10 mg/L polysulfone colloids added to 150 mL of 15 mg/L humic acid solution. Humic acid adsorption was performed in accord with the method described in Example 2. The humic acid was then desorbed from the colloids by raising the pH of the solution to 12 with 1N sodium hydroxide.

[0040] Readsorption of humic acid was performed by adjusting the pH of the solution down to 6.5 with 1N HCl. Since most humic acid adsorption occurred on the first day, the experiments were performed with 1 day adsorption, 1 day desorption, and 1 day readsorption time periods.

Example 4

[0041] The adsorption capacity of polysulfone colloids was compared to activated carbon as follows: A set of adsorption isotherms were performed with humic acid solution on activated carbon with carbon doses of 0, 5, 10, 20, 50, and 100 mg of activated carbon per Liter of humic acid solution.

[0042] NORIT SA-UF activated carbon, obtained from NORIT Americas, Atlanta, GA, was used because it has the best adsorption capacity among the commercialized activated carbons. The initial concentration of humic acid was 5 mg/L and the solution pH was adjusted to 6.5. The experiment lasted 7 days and was considered to be in equilibrium on the 7th day.

[0043] Adsorption isotherms of humic acid on activated carbon versus polysulfone colloids are shown in Figure 2. For activated carbon adsorption, the linear fit of the first three data excluding the curvature gives a $F = 0.072$, and $1/n = 0.7695$ with Freundlich isotherm fitting. Compared to the activated carbon, the polymer colloids demonstrated greater adsorption capacity, approaching an order of magnitude, as shown by the upper line.

Prophetic Example 5

[0044] Colloids are incorporated into a membrane by vacuum filtering a water suspension containing the colloids through a 0.22 micron pore diameter nylon filter. Even though the pore diameter of the filter is larger than the approximate 50 nm diameter of the colloid particles, the particles are trapped due to aggregation. Water contaminated with organic matter is then purified of organic matter by passing the contaminated water through the filter.